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Full worked solutions in: Chemsheets AS029 (Calorimetry) TASK 2 – Calorimetry calculations 2 1  
-3800 kJ mol<sup>-1</sup> 2 -51.8 kJ mol<sup>-1</sup> 3 -2500 kJ mol<sup>-1</sup> 4 +20.9 kJ mol<sup>-1</sup> 5 -4203 kJ mol<sup>-1</sup> 6 -3920 kJ mol<sup>-1</sup>

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7 -54.3 kJ mol<sup>-1</sup> 8 -1145 kJ mol<sup>-1</sup> 9 -1223 kJ mol<sup>-1</sup> 10 -154 kJ mol<sup>-1</sup> 11 exo, exo, endo 12 -503 kJ mol<sup>-1</sup>

### **Chemsheets AS 029 (Thermodynamics) ANS 2**

2O<sub>4</sub> 2-) Example complex N O [Cr(NH<sub>2</sub>CH<sub>2</sub>CH<sub>2</sub>NH<sub>2</sub>)<sub>3</sub>]<sup>3+</sup> [Cr(C<sub>2</sub>O<sub>4</sub>)<sub>3</sub>]<sup>3-</sup> Multidentate ligands - ligands which form more than two co-ordinate bonds to a metal ion Ligand EDTA<sup>4-</sup> forms 6 bonds porphyrin forms 4 bonds How it bonds Example [Cu(EDTA)]<sup>2-</sup> haemoglobin CH<sub>2</sub> CH<sub>2</sub> H<sub>2</sub>N: :NH<sub>2</sub>-O: C:O-O O NH<sub>2</sub> Cr NH<sub>2</sub> H<sub>2</sub>N NH<sub>2</sub> H<sub>2</sub>NH<sub>2</sub> 3+ O Cr NH<sub>2</sub> O O O O C ...

### **Chemsheets A2 1027 (Transition Metals)**

Example 3 50.0 cm<sup>3</sup> of 0.500 mol dm<sup>-3</sup> copper sulphate solution was put in a calorimeter and 2.0 g of zinc powder added. The temperature of the solution rose by 24.0 C.

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© **www.CHEMSHEETS.co.uk 23-Feb-2016 Chemsheets A2 1001 Page 2 TASK 1 - Finding rates using gradients**

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2,2-dimethylamine 2-methylbut-1-ene 3-methylbutan-1-ol 4-hydroxybutanone 3-methylbutanoic acid butyl methanoate ethanenitrile propanal Decide which spectrum belongs to which compound and draw the molecule next to the spectrum.

### **Chemsheets A2 029 (Spectroscopy) - Chemistry Home page**

n is called the order of reaction with respect to that reactant, and is usually 2, 1 or 0 Order Zero order First order Second order n 0 1 2 The effect the reaction rate is proportional the reaction rate is proportional to [A]<sup>0</sup> (= 1) and so the rate is not affected by changes in [ A] to [A] to [A]<sup>2</sup> [A] x 2

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rate unchanged rate x 22 (i.e. x 4)

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4 3 2 1 0

### **SECTION 1 1H NMR**

the value of  $-114.2 \text{ kJ mol}^{-1}$  in the first equation means that  $114.2 \text{ kJ}$  of heat energy is released when 1 mole of  $\text{H}_2\text{SO}_4$  reacts with 2 moles of  $\text{NaOH}$ . the value of  $-57.1 \text{ kJ mol}^{-1}$  in the second equation means that  $57.1 \text{ kJ}$  of heat energy is released when  $\frac{1}{2}$  mole of  $\text{H}_2\text{SO}_4$  reacts with 1 mole of  $\text{NaOH}$ .

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Some gas phase examples. [www.chemsheets.co.uk](http://www.chemsheets.co.uk) A2 1029 01-Jun-2016 HOMOGENEOUS CATALYSTS. 1) Catalyst reacts with a reactant to form intermediate. 2) Intermediate reacts to form product faster than original reactant (and regenerates catalyst). e.g. acid catalyst  $\text{X} + \text{Y}$  products.  
1)  $\text{X} + \text{H}^+ \dots$

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