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Roniyah2002. Formula:  $Q = mc\Delta T$ . Key  
Concepts: Terms in this set (9) Heat  
Energy (Q): 63,536. If 200 grams of  
water is to be heated from  $24.0^{\circ}\text{C}$  to  
 $100^{\circ}\text{C}$  to make a cup of tea, how much  
heat must be added? The specific heat  
of water ...

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## **Specific Heat Practice Problems Flashcards | Quizlet**

Solution: Use the formula  $q = mc\Delta T$   
where  $q$  = heat energy  $m$  = mass  $c$  =  
specific heat  $\Delta T$  = change in  
temperature Putting the numbers into  
the equation yields:  $487.5 \text{ J} = (25 \text{ g})c(75$   
 $^{\circ}\text{C} - 25 ^{\circ}\text{C})$   $487.5 \text{ J} = (25 \text{ g})c(50 ^{\circ}\text{C})$

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Solve for  $c$ :  $c = 487.5 \text{ J}/(25\text{g})(50 \text{ }^\circ\text{C})$   $c = 0.39 \text{ J/g}\cdot^\circ\text{C}$

## **Specific Heat Worked Example Problem - ThoughtCo**

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Janus\_Han. Formula:  $Q = mc\Delta T$ . Terms in

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this set (9) If 200 grams of water is to be heated from  $24.0^{\circ}\text{C}$  to  $100^{\circ}\text{C}$  to make a cup of tea, how much heat must be added? The specific heat of water is  $4.18 \text{ J/g}^{\circ}\text{C}$ .

## **Specific Heat Practice Problems Flashcards | Quizlet**

Specific Heat Practice Problems  $Q = mc$

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Specific Heat of liquid water =  $4.184 \text{ J/g } ^\circ\text{C}$   
 $1 \text{ cal/g } ^\circ\text{C} = 4.184 \text{ J/g } ^\circ\text{C}$   
 $1 \text{ food Calorie} = 1000 \text{ calories}$   
 $1 \text{ calorie} = 4.184 \text{ J}$   
Q water Q rxn 1. A slice of pizza contains 180 nutritional Calories. Convert this energy into joules. 2.

## **Specific Heat Practice Problems - FCPS**

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from 25°C to 115°C. Find the specific heat of aluminum. 7) The specific heat of lead (Pb) is 0.129 J/g °C. Find the amount of heat released when 2.4 mol of lead are cooled from 37.2°C to 22.5°C. ADVANCED CALORIMETRY 8) If 150.0 grams of iron at 95.0 °C, is placed in an insulated container containing 500.0 grams of

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**) ( $\Delta T$**

Specific Heat Problems. Specific Heat Problems. 1) How much heat must be absorbed by 375 grams of water to raise its temperature by  $25^{\circ}\text{C}$ ? 2) What mass of water can be heated from  $25.0^{\circ}\text{C}$  to  $50.0^{\circ}\text{C}$  by the addition of 2825 J? 3) What is the final temperature when 625

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grams of water at 75.0° C loses 7.96 x 104J? 4) A copper cylinder has a mass of 76.8 g and a specific heat of 0.092 cal/g·C.

## **Specific Heat Problems - mmsphyschem.com**

HEAT Practice Problems .  $Q = m \times \Delta T \times C$   
. 5.0 g of copper was heated from 20°C

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to  $80^{\circ}\text{C}$ . How much energy was used to heat Cu? (Specific heat capacity of Cu is  $0.092 \text{ cal/g } ^{\circ}\text{C}$ ) 27.6 cal. How much heat is absorbed by 20g granite boulder as energy from the sun causes its temperature to change from  $10^{\circ}\text{C}$  to  $29^{\circ}\text{C}$ ? (Specific heat capacity of granite is  $0.1 \text{ cal/g}^{\circ}\text{C}$ ) 38 cal

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## Answers

### **HEAT Practice Problems**

Problem #4: A 35.0 g block of metal at 80.0 °C is added to a mixture of 100.0 g of water and 15.0 g of ice in an isolated container. All the ice melted and the temperature in the container rose to 10.0 °C. What is the specific heat of the metal? Solution: 1) Determine heat required to melt the ice:

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## **ChemTeam: How to Determine Specific Heat: Problem 1 - 10**

Worksheet- Calculations involving  
Specific Heat 1. For  $q = m c \Delta T$  : identify  
each variables by name & the units  
associated with it.  $q$  = amount of heat (J)  
 $m$  = mass (grams)  $c$  = specific heat  
(J/g°C)  $\Delta T$  = change in temperature (°C)

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2. Heat is not the same as temperature, yet they are related. Explain how they differ from each other.

## **Worksheet- Calculations involving Specific Heat**

As you can see, many problems mix the concepts of specific heat and latent heat.

$$2) \text{ b) } Q = L f m = 3.33 \times 10^5 \times 1 = 3.33$$

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$\times 10^5 \text{ J}$ . In this case,  $Q_1 < Q_2$  so that all the ice will be melted. What will be the final temperature of the mixture?

## **Calculations involving specific heat, heat and latent heat ...**

By comparison, look at the heat capacity of copper. 1 gram of copper will rise in temperature by  $1 \text{ C}^\circ$  when just 0.385

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Joules of heat is absorbed. This low specific heat capacity indicates that copper is a good conductor of heat. You might predict that applying a small amount of heat will make the temperature of a gram of copper skyrocket while the same amount of heat hardly makes the ...

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## **Chemistry: Specific Heat Capacity - AlgebraLAB**

<https://getchemistryhelp.com/learn-chemistry-fast/> These example problems show you how to solve for heat, specific heat, and the change in temperature using ...

## **Chemistry Practice Problems: Heat**

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## Answers **and Specific Heat - YouTube**

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Theory Calculus Probability ... Specific  
heat Phase changes Challenge Quizzes  
Phase transitions: Level 2-4 Challenges  
Specific heat . Consider an aluminium  
cup with mass 140.0 g 140.0 ... Problem  
Loading...

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Answers

## **Specific heat Practice Problems Online | Brilliant**

This is a conservation of energy problem. The heat gained by the ice will be equal to the heat lost by the coffee.  
 $+Q_{\text{ice}} = -Q_{\text{coffee}}$  This mixing problem is more complicated than the ones in the previous section, however.

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Answers

## **Latent Heat - Practice - The Physics Hypertextbook**

This chemistry tutorial covers the difference between heat capacity and specific heat and includes several examples of how to find specific heat and how to u...

## **Heat Capacity and Specific Heat -**

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## Answers

### **Chemistry Tutorial - YouTube**

Specific Heat Capacity Practice Problems

Name: 1. When 3.0 kg of water is cooled from 80.0 C to 10.0 C, how much heat energy is lost? 2. How much heat is needed to raise a 0.30 kg piece of aluminum from 30. C to 150 C? 3. Calculate the temperature change when:  
a) 10.0 kg of water loses 232 kJ of heat.

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## **Specific Heat Capacity Practice Problems - Studylib**

[View the accompanying Heat & Specific Heat Capacity Practice Problems here.]

Temperature vs. Heat Temperature - The average energy of individual particles in motion. For example, the temperature of a cup of coffee is the

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average energy of all of the ... Read  
More

## **Chemistry Lesson: Heat & Specific Heat Capacity - Get ...**

The specific heat capacity of water is  $4200 \text{ J/kg}^\circ\text{C}$ . Rotate to landscape screen format on a mobile phone or small tablet to use the Mathway widget, a free math

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problem solver that answers your questions with step-by-step explanations

- .

## **Specific Heat Capacity (examples, solutions, videos, notes)**

Specific Heat Capacity (C or S )-The quantity of heat required to raise the temperature of a substance by one

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degree Celsius is called the specific heat capacity of the substance. The quantity of heat is frequently measured in units of Joules(J). Another property, the specific heat, is the heat capacity of the substance per gram of the substance.

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